



GCSE
COMBINED SCIENCE: TRILOGY
8464/C/1F

Chemistry Paper 1F

Mark scheme

June 2022

Version: 1.0 Final Mark Scheme



2 2 6 G 8 4 6 4 C 1 F / M S

Mark schemes are prepared by the Lead Assessment Writer and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation events which all associates participate in and is the scheme which was used by them in this examination. The standardisation process ensures that the mark scheme covers the students' responses to questions and that every associate understands and applies it in the same correct way. As preparation for standardisation each associate analyses a number of students' scripts. Alternative answers not already covered by the mark scheme are discussed and legislated for. If, after the standardisation process, associates encounter unusual answers which have not been raised they are required to refer these to the Lead Examiner.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of students' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

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Information to Examiners

1. General

The mark scheme for each question shows:

- the marks available for each part of the question
- the total marks available for the question
- the typical answer or answers which are expected
- extra information to help the examiner make their judgement
- the Assessment Objectives and specification content that each question is intended to cover.

The extra information is aligned to the appropriate answer in the left-hand part of the mark scheme and should only be applied to that item in the mark scheme.

At the beginning of a part of a question a reminder may be given, for example: where consequential marking needs to be considered in a calculation; or the answer may be on the diagram or at a different place on the script.

In general the right-hand side of the mark scheme is there to provide those extra details which confuse the main part of the mark scheme yet may be helpful in ensuring that marking is straightforward and consistent (for example, a scientifically correct answer that could not reasonably be expected from a student's knowledge of the specification).

2. Emboldening and underlining

- 2.1** In a list of acceptable answers where more than one mark is available 'any **two** from' is used, with the number of marks emboldened. Each of the following bullet points is a potential mark.
- 2.2** A bold **and** is used to indicate that both parts of the answer are required to award the mark.
- 2.3** Alternative answers acceptable for a mark are indicated by the use of **or**.
Alternative words in the mark scheme are shown by a solidus eg allow smooth / free movement.
- 2.4** Any wording that is underlined is essential for the marking point to be awarded.

3. Marking points

3.1 Marking of lists

This applies to questions requiring a set number of responses, but for which students have provided extra responses. The general principle to be followed in such a situation is that 'right + wrong = wrong'.

Each error / contradiction negates each correct response. So, if the number of errors / contradictions equals or exceeds the number of marks available for the question, no marks can be awarded.

However, responses considered to be neutral (indicated as * in example 1) are not penalised.

Example 1: What is the pH of an acidic solution?

[1 mark]

Student	Response	Marks awarded
1	green, 5	0
2	red*, 5	1
3	red*, 8	0

Example 2: Name **two** magnetic materials.

[2 marks]

Student	Response	Marks awarded
1	iron, steel, tin	1
2	cobalt, nickel, nail*	2

3.2 Use of symbols / formulae

If a student writes a chemical symbol / formula instead of a required chemical name, or uses symbols to denote quantities in a physics equation, full credit can be given if the symbol / formula is correct and if, in the context of the question, such action is appropriate.

3.3 Marking procedure for calculations

Marks should be awarded for each stage of the calculation completed correctly, as students are instructed to show their working. At any point in a calculation students may omit steps from their working. If a subsequent step is given correctly, the relevant marks may be awarded.

Full marks are **not** awarded for a correct final answer from incorrect working.

3.4 Interpretation of 'it'

Answers using the word 'it' should be given credit only if it is clear that the 'it' refers to the correct subject.

3.5 Errors carried forward

An error can be carried forward from one question part to the next and is shown by the abbreviation 'ecf'.

Within an individual question part, an incorrect value in one step of a calculation does not prevent all of the subsequent marks being awarded.

3.6 Phonetic spelling

Marks should be awarded if spelling is not correct but the intention is clear, **unless** there is a possible confusion with another technical term.

3.7 Brackets

(.....) are used to indicate information which is not essential for the mark to be awarded but is included to help the examiner identify the sense of the answer required.

3.8 Allow

In the mark scheme additional information, 'allow' is used to indicate creditworthy alternative answers.

3.9 Ignore

Ignore is used when the information given is irrelevant to the question or not enough to gain the marking point. Any further correct amplification could gain the marking point.

3.10 Do **not** accept

Do **not** accept means that this is a wrong answer which, even if the correct answer is given as well, will still mean that the mark is not awarded.

3.11 Numbered answer lines

Numbered lines on the question paper are intended to support the student to give the correct number of responses. The answer should still be marked as a whole.

4. Level of response marking instructions

Extended response questions are marked on level of response mark schemes.

- Level of response mark schemes are broken down into levels, each of which has a descriptor.
- The descriptor for the level shows the average performance for the level.
- There are two marks in each level.

Before you apply the mark scheme to a student's answer, read through the answer and, if necessary, annotate it (as instructed) to show the qualities that are being looked for. You can then apply the mark scheme.

Step 1: Determine a level

Start at the lowest level of the mark scheme and use it as a ladder to see whether the answer meets the descriptor for that level.

The descriptor for the level indicates the different qualities that might be seen in the student's answer for that level. If it meets the lowest level then go to the next one and decide if it meets this level, and so on, until you have a match between the level descriptor and the answer. With practice and familiarity you will find that for better answers you will be able to quickly skip through the lower levels of the mark scheme.

When assigning a level you should look at the overall quality of the answer. Do **not** look to penalise small and specific parts of the answer where the student has not performed quite as well as the rest. If the answer covers different aspects of different levels of the mark scheme you should use a best fit approach for defining the level.

Use the variability of the response to help decide the mark within the level, ie if the response is predominantly level 2 with a small amount of level 3 material it would be placed in level 2 but be awarded a mark near the top of the level because of the level 3 content.

Step 2: Determine a mark

Once you have assigned a level you need to decide on the mark. The descriptors on how to allocate marks can help with this. The exemplar materials used during standardisation will help. There will be an answer in the standardising materials which will correspond with each level of the mark scheme. This answer will have been awarded a mark by the Lead Examiner. You can compare the student's answer with the example to determine if it is the same standard, better or worse than the example. You can then use this to allocate a mark for the answer based on the Lead Examiner's mark on the example.

You may well need to read back through the answer as you apply the mark scheme to clarify points and assure yourself that the level and the mark are appropriate.

Indicative content in the mark scheme is provided as a guide for examiners. It is not intended to be exhaustive and you must credit other valid points. Students do not have to cover all of the points mentioned in the indicative content to reach the highest level of the mark scheme.

You should ignore any irrelevant points made. However, full marks can be awarded only if there are no incorrect statements that contradict a correct response.

An answer which contains nothing of relevance to the question must be awarded no marks.

Question 1

Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.1	alkali metals		1	AO1 5.1.2.5

Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.2	3 / three	must be in this order	1	AO2 5.1.1.5
	4 / four		1	

Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.3	+1		1	AO2 5.1.1.4

Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.4	<p>Stage</p> <p>Stage 1</p> <p>Stage 2</p> <p>do not accept more than one line from a box on the left</p>	<p>Physical change</p> <p>boiling</p> <p>condensing</p> <p>dissolving</p> <p>freezing</p> <p>melting</p>	<p>1</p> <p>1</p>	AO1 5.2.2.1

Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.5	28 (°C)		1	AO2 5.1.2.5

Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.6	bar plotted at 63 °C	allow a tolerance of $\pm \frac{1}{2}$ a small square	1	AO3 5.1.2.5

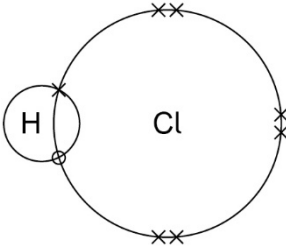
Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.7	(melting point) decreases as go down Group 1	allow converse in terms of increasing melting point ignore boiling points	2	AO1
	(and the) differences get smaller or initial drop is largest	allow 1 mark for (melting point) decreases	1	AO3 5.1.2.5

Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.8	liquid		1	AO3 5.1.2.5 5.2.2.1

Question	Answers	Extra information	Mark	AO / Spec. Ref.
01.9	loses an electron		1	AO1 5.1.1.7 5.1.2.1 5.1.2.5

Total Question 1		13
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Question 2

Question	Answers	Extra information	Mark	AO / Spec. Ref.
02.1	one shared pair in overlap	allow any combination of circles, dots, crosses or e ⁽⁻⁾	1	AO1 5.1.2.6 5.2.1.4
	6 non-bonding electrons in outer shell of chlorine	do not accept extra electron(s) on outer shell of hydrogen ignore any inner shell electrons an answer of  scores 2 marks	1	

Question	Answers	Extra information	Mark	AO / Spec. Ref.
02.2	sodium chloride	allow NaCl	1	AO2 5.4.2.2

Question	Answers	Extra information	Mark	AO / Spec. Ref.
02.3	5.85 (g)		1	AO2 5.3.1.1 5.4.2.2 5.5.1.1 RPA10

Question	Answers	Extra information	Mark	AO / Spec. Ref.
02.4	17.6 (°C)		1	AO2 5.4.2.2 5.5.1.1 RPA10

Question	Answers	Extra information	Mark	AO / Spec. Ref.
02.5	exothermic (reaction)	allow neutralisation (reaction)	1	AO2 5.5.1.1 RPA10

Question	Answers	Extra information	Mark	AO / Spec. Ref.
02.6	Na ⁺		1	AO1 AO2 5.4.2.2 5.4.2.4 5.5.1.1
	OH ⁻		1	

Total Question 2		8
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Question 3

Question	Answers	Extra information	Mark	AO / Spec. Ref.
03.1	giant covalent		1	AO1 5.2.2.6 5.2.3.1

Question	Answers	Extra information	Mark	AO / Spec. Ref.
03.2	4 / four		1	AO3 5.2.2.6 5.2.3.1

Question	Answers	Extra information	Mark	AO / Spec. Ref.
03.3	hexagons		1	AO1 5.2.3.3

Question	Answers	Extra information	Mark	AO / Spec. Ref.
03.4	carbon		1	AO1 5.2.3.3

Question	Answers	Extra information	Mark	AO / Spec. Ref.
03.5	electronics		1	AO1 5.2.3.3

Question	Answers	Extra information	Mark	AO / Spec. Ref.
03.6	covalent bonds		1	AO1 5.2.2.5

Question	Answers	Extra information	Mark	AO / Spec. Ref.
03.7	intermolecular		1	AO1 5.2.2.5

Question	Answers	Extra information	Mark	AO / Spec. Ref.
03.8	(magnesium) 22 and (copper) 3	allow correct use of incorrectly determined value(s) for number of magnesium atoms and / or copper atoms	1	AO2 5.2.2.7
	(percentage =) $\frac{3}{25} (\times 100)$		1	
	=12 (%)		1	

Question	Answers	Extra information	Mark	AO / Spec. Ref.
<p>03.9</p>	<p>(alloy is harder because) copper atoms are larger or (copper) atoms are a different size</p>	<p>allow (so) the atoms cannot slide over each other</p>	1	AO2
	<p>(so the) layers of (magnesium) atoms are distorted</p>		1	AO2
	<p>(and therefore the) layers cannot easily slide</p>		1	AO1 5.2.1.5 5.2.2.7

Total Question 3		13
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Question 4

Question	Answers	Extra information	Mark	AO / Spec. Ref.
04.1	$2 \text{Mg} + \text{O}_2 \rightarrow 2 \text{MgO}$	allow multiples	1	AO2 5.1.2.5 5.4.1.1

Question	Answers	Extra information	Mark	AO / Spec. Ref.
04.2	any one from: <ul style="list-style-type: none"> • wear safety glasses / goggles • do not look directly at burning magnesium • wear heat proof glove 	allow look through blue glass allow use tongs allow tie hair back	1	AO3 5.4.1.1 5.4.1.2 5.5.1.1

Question	Answers	Extra information	Mark	AO / Spec. Ref.
04.3	$(M_r =) 24 + (2 \times 19)$ $= 62$		1 1	AO2 5.1.2.5 5.3.1.2

Question	Answers	Extra information	Mark	AO / Spec. Ref.
04.4	any one from: <ul style="list-style-type: none"> • argon has a full outer shell • argon has 8 electrons in the outer shell • argon has a stable arrangement of electrons. 	allow energy level for shell	1	AO2
	(so) argon is unreactive	allow does not need to lose and / or gain electrons ignore argon is a noble gas	1	AO1 5.1.2.4

Question	Answers	Extra information	Mark	AO / Spec. Ref.										
04.5	<table border="0"> <thead> <tr> <th style="text-align: center;">Metal</th> <th style="text-align: center;">Method used to extract that metal</th> </tr> </thead> <tbody> <tr> <td style="border: 1px solid black; text-align: center; width: 150px; height: 25px;">Metal D</td> <td style="border: 1px solid black; padding: 5px;">Extracted by electrolysis of a molten ionic compound.</td> </tr> <tr> <td style="border: 1px solid black; text-align: center; width: 150px; height: 25px;">Metal E</td> <td style="border: 1px solid black; padding: 5px;">Extracted from its oxide by reduction with carbon.</td> </tr> <tr> <td></td> <td style="border: 1px solid black; padding: 5px;">Extracted from its oxide by reduction with hydrogen.</td> </tr> <tr> <td></td> <td style="border: 1px solid black; padding: 5px;">Removed from the Earth as the metal itself.</td> </tr> </tbody> </table>	Metal	Method used to extract that metal	Metal D	Extracted by electrolysis of a molten ionic compound.	Metal E	Extracted from its oxide by reduction with carbon.		Extracted from its oxide by reduction with hydrogen.		Removed from the Earth as the metal itself.		1	AO3 5.4.1.2 5.4.1.3 5.4.3.1 5.4.3.2 5.4.3.3
		Metal	Method used to extract that metal											
Metal D	Extracted by electrolysis of a molten ionic compound.													
Metal E	Extracted from its oxide by reduction with carbon.													
	Extracted from its oxide by reduction with hydrogen.													
	Removed from the Earth as the metal itself.													
	do not accept more than one line from a box on the left		1											

Question	Answers	Extra information	Mark	AO / Spec. Ref.
04.6	(because) <u>ions</u>	allow reference to charged particles for MP2 and MP3	1	AO1 5.2.2.3
	(which are) free to move when molten		1	
	(but are) fixed in solid		1	

Question	Answers	Extra information	Mark	AO / Spec. Ref.
04.7	(sodium contains) delocalised electrons	allow free electrons	1	AO1 5.2.2.8
	(which) carry (electrical) charge (through the metal / sodium)	ignore throughout for through ignore current / electricity	1	

Total Question 4		13
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Question 5

Question	Answers	Extra information	Mark	AO / Spec. Ref.
05.1	any three from: <ul style="list-style-type: none"> • green solid / powder • colourless solution • blue solution formed • copper carbonate disappears • fizzing / effervescence or bubbles (of gas) • stops fizzing • solid / powder left at the end or copper carbonate left at the end 	ignore green copper carbonate allow colour (of solution) changes allow solid disappears ignore gas allow fizzing slows down allow (container) gets hot or allow temperature increases	3	AO2 AO3 5.4.2.2 5.4.2.3 RPA8

Question	Answers	Extra information	Mark	AO / Spec. Ref.
05.2	filtration or filter		1	AO1 5.4.2.3 RPA8

Question	Answers	Extra information	Mark	AO / Spec. Ref.
05.3	7		1	AO1 5.4.2.3 5.4.2.4 RPA8

Question	Answers	Extra information	Mark	AO / Spec. Ref.
05.4	neutralisation	allow exothermic	1	AO1 5.4.2.2 5.4.2.4 RPA8

Question	Answers	Extra information	Mark	AO / Spec. Ref.
05.5	83 (g at 80 °C)	allow a value in range 82–84 (g at 80 °C)	1	AO2
	32 (g at 20 °C)	allow a value in range 32–33 (g at 20 °C)	1	AO2
	(83–32 =) 51 (g)	allow a correct calculation using incorrectly read values for mass at 80 °C and/or 20 °C	1	AO3 5.4.2.3

Total Question 5		9
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Question 6

Question	Answers	Extra information	Mark	AO / Spec. Ref.
06.1	K_2SO_4		1	AO2 5.1.1.1 5.4.3.4 RPA9

Question	Answers	Extra information	Mark	AO / Spec. Ref.
06.2	(volume of hydrogen) 30 (cm ³) and (volume of oxygen) 15 (cm ³)		1	AO2 5.4.3.4 RPA9

Question	Answers	Extra information	Mark	AO / Spec. Ref.
06.3	(because) the ratio of volume of hydrogen : oxygen is 2 : 1		1	AO3 5.4.3.4 RPA9
	(and this is the) same as the ratio of hydrogen (atoms) : oxygen (atoms) in (formula of) H ₂ O		1	
	OR	must relate to the volumes given in question 06.2		
	(because) the ratio of volume of hydrogen : oxygen is not 2 : 1 (1)			
	(and this is) different to the ratio of hydrogen (atoms) : oxygen (atoms) in (formula of) H ₂ O (1)			

Question	Answers	Extra information	Mark	AO / Spec. Ref.
06.4	$9 \pm 3 \text{ cm}^3$		1	AO2 5.3.1.4

Question	Answers	Extra information	Mark	AO / Spec. Ref.
06.5	(conversion) $\left(\frac{25}{1000} = \right) 0.025 \text{ (dm}^3\text{)}$		1	AO2 5.3.2.5 5.4.3.4
	(concentration =) $\frac{0.86}{0.025}$	allow correct use of incorrect / no conversion	1	
	= 34.4 (g per dm ³)	allow 34 (g per dm ³)	1	
	OR			
	(conversion) $\frac{1000}{25} \text{ (1)}$			
	= 40 (1)			
(40 × 0.86) = 34.4 (g per dm ³) (1)	allow correct use of incorrect / no conversion allow 34 (g per dm ³)			
OR				
(concentration =) $\frac{0.86}{25} \text{ (1)}$				
= 0.0344 (1)				
(conversion) (0.0344 × 1000) = 34.4 (g per dm ³) (1)	allow 34 (g per dm ³)			

Total Question 6		8
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Question 7

Question	Answers	Mark	AO / Spec. Ref.
07	Level 3: The method would lead to the production of a valid outcome. All key steps are identified and logically sequenced.	5–6	AO3
	Level 2: The method would not necessarily lead to a valid outcome. Most steps are identified, but the method is not fully logically sequenced.	3–4	AO3
	Level 1: The method would not lead to a valid outcome. Some relevant steps are identified, but links are not made clear.	1–2	AO1
	No relevant content.	0	
	<p>Indicative Content</p> <ul style="list-style-type: none"> • measure volume of (hydrochloric) acid • into a suitable container eg polystyrene cup • measure the initial temperature (of hydrochloric acid) • with a thermometer • add stated mass of one metal • stir <ul style="list-style-type: none"> • measure the highest temperature reached of the solution or measure temperature reached after a set time period • determine the temperature difference <ul style="list-style-type: none"> • repeat <ul style="list-style-type: none"> • repeat for each metal • with same mass • in same physical state (powder, lump, etc) • with the same volume and / or concentration of (hydrochloric) acid <ul style="list-style-type: none"> • use results to arrange metals in order of reactivity • most reactive metal has the largest temperature change <p>to access level 3 there must be an indication of how the temperature change is determined with the same mass of the 3 different metals reacted with the same volume of (hydrochloric) acid</p>		5.1.4.2 RPA2

Total Question 7		6
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